An efficient green synthesis of polyfunctional pyrazole-triazole hybrids and bis-triazoles via chromium incorporated fluorapatite encapsulated iron oxide nanocatalyst

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1. Introduction

Nitrogen containing heterocyclic rings are momentous groups in organic chemistry. Derivatives of triazole rings are an essential aromatic five-membered heterocycles presenting important biological activities. Some of these compounds have shown anti-bacterial, anti-fungal, anti-inflammatory, anti-tumor, antimalarial and anti-cancer properties. Numerous derivatives of triazole rings are widespread in natural product and pharmacological compounds. Some representative examples are presented in Fig. 1. Therefore, various methods and catalysts have been introduced for the synthesis of triazole derivatives such as, application of a microwave-assisted click chemistry using copper(I), samarium doped fluorapatites, [C16MPy]AlCl3Br as ionic liquid and potassium hydroxide. Nevertheless, the most general method for the synthesis of five-membered heterocycles is the cycloaddition reaction which has been recently discussed in detail. Some of the reported procedures suffer from harsh reaction conditions, complex synthetic pathways and non-
recyclability of the catalyst. Consequently, more facile synthetic methods are still required. To achieve this objective, we developed an efficient protocol by using the magnetic chromium incorporated fluoroapatite ($\gamma$-Fe$_2$O$_3$@FAp@Cr) as a novel catalyst for the green synthesis of versatile polyfunctional pyrazole-triazole hybrids in short reaction time (10-60 min) and excellent yield (88-95%). Interestingly, the method was also extended to the synthesis of bis-triazoles successfully.

2. Results and Discussion

Following our continued studies in the benign synthesis of biologically important heterocycles$^{31-39}$, we have developed a convenient method for the efficient synthesis of mono- and bis-triazole in the presence of newly synthesized magnetic nanocatalyst ($\gamma$-Fe$_2$O$_3$@FAp@Cr). Initially, the requisite $\gamma$-Fe$_2$O$_3$@FAp was prepared according to the literature report$^{37,38}$ and reacted with CrCl$_3$.6H$_2$O in water at room temperature to furnish the desired catalyst (Scheme 1). The structure of the catalyst was established by FT-IR, XRD, SEM, EDX and TEM.

FT-IR analysis

In the FT-IR spectra of $\gamma$-Fe$_2$O$_3$@FAp@Cr NPs the bending vibrations of P-O-P which are overlapping with the stretching vibration of Fe-O are visible at 589 and 604 cm$^{-1}$. The stretching vibrations of P-O bands appeared at 1039 cm$^{-1}$. The broad and strong band at 3415 cm$^{-1}$ belongs to the stretching vibrations of O-H groups and absorbed water (Fig. 2).

XRD analysis

Fig. 3 shows the XRD analysis of the $\gamma$-Fe$_2$O$_3$@FAp@Cr catalyst in contrast to $\gamma$-Fe$_2$O$_3$ and FAp. This pattern shows characteristic peaks at around $2\theta = 35.7^\circ$, $48.7^\circ$, $52.4^\circ$, $53.4^\circ$, $56.2^\circ$, $63.3^\circ$ which are readily distinguished from the XRD pattern. They agree with the cubic structure of maghemite (JCPDS file No. 39–1346). Diffraction peaks at around $2\theta = 26.1^\circ$, $28.2^\circ$, $29.2^\circ$, $32.0^\circ$, $33.2^\circ$, $34.3^\circ$, $40.1^\circ$, $46.9^\circ$, $49.7^\circ$, $51.7^\circ$, $78.7^\circ$ are related to the FAp (JCPDS File No. 003-9137). The average crystallite size was calculated to be 25 nm for $\gamma$-Fe$_2$O$_3$@FAp@Cr using the Scherrer equation.
**Scheme 1.** Synthesis of γ-Fe₂O₃@FAp@Cr NPs.

**Fig. 2.** The FT-IR spectra of γ-Fe₂O₃@FAp@Cr
Scanning electron microscopy analysis (SEM)

The morphology and particle size of the γ-Fe2O3@FAp@Cr catalyst were investigated using SEM technique (Fig. 4). According to the SEM images γ-Fe2O3@FAp@Cr MNPs are formed with almost spherical morphology. The average size of γ-Fe2O3@FAp@Cr nanoparticles is about 15-40 nm according to the measurement software.

EDX analysis

The results of energy dispersive X-ray spectroscopy (EDX) analysis of the synthesized γ-Fe2O3@FAp@Cr MNPs proved existence of Fe (24.0 w/w%), O (48.3 w/w%), P (7.9 w/w%) Ca (15.0 w/w%), F (0.7 w/w%) and Cr (4.1 w/w%) atoms in the structure that confirms the presence of γ-Fe2O3 core in the structure of MNPs (Fig. 5).

TEM analysis

The morphology and size of the γ-Fe2O3@FAp@Cr MNPs were checked by the TEM spectrum as shown in Fig. 6. According to the TEM images analysis, the size of these nanoparticles was estimated 15-25 nm.

In order to study the catalytic capability of the synthesized γ-Fe2O3@FAp@Cr nanoparticles in organic reactions, we decided to investigate its activity in a green synthesis of several pyrazole-triazole hybrids (Scheme 1). Therefore, for optimization of the reaction conditions as a model reaction, thiosmicarbazide or semicarbazide (1) and 1,3-diphenyl-1H-pyrazole-4-carbaldehyde (2) in the presence of γ-Fe2O3@FAp@Cr nanoparticles in a variety of solvents and various temperatures were reacted (Table 1).
Scheme 1. Synthesis of 5-(3-(aryl)-1-phenyl-1H-pyrazol-4-yl)-1,2,4-triazolidin-3-ones (thiones) derivatives via γ-Fe2O3@FAp@Cr.

It is evident from the results that using chloroform at room temperature leads to the desired product 5-(1,3-diphenyl-1H-pyrazol-4-yl)-1,2,4-triazolidin-3-thione (3a) in 10 min and 95% yield (Table 1, Entry 4). To demonstrate the efficiency of the nanocatalyst the reaction was performed in the absence of the catalyst and in the presence of various acidic and basic catalysts (Table 2). This study revealed that the reaction in the presence of γ-Fe2O3@FAp@Cr nanocatalyst produces better result. The amount of the catalyst was also verified which proved that the use of 0.06 gr (4.7 mol%) of the catalyst per mmol substrate provides the best yield of triazolidin-3-thione (3a).

Table 1. Synthesis of 3a in the presence of γ-Fe2O3@FAp@Cr in various solvents and temperatures

<table>
<thead>
<tr>
<th>Entry</th>
<th>Solvent</th>
<th>Temperature (°C)</th>
<th>Time (min)</th>
<th>Yield (%)a,b</th>
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<tbody>
<tr>
<td>1</td>
<td>Toluene</td>
<td>25</td>
<td>120</td>
<td>trace</td>
</tr>
<tr>
<td>2</td>
<td>THF</td>
<td>25</td>
<td>120</td>
<td>65</td>
</tr>
<tr>
<td>3</td>
<td>DMF</td>
<td>25</td>
<td>120</td>
<td>70</td>
</tr>
<tr>
<td>4</td>
<td>CHCl3</td>
<td>25</td>
<td>10</td>
<td>95</td>
</tr>
<tr>
<td>5</td>
<td>CHCl3</td>
<td>40</td>
<td>10</td>
<td>95</td>
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<tr>
<td>6</td>
<td>CHCl3</td>
<td>60</td>
<td>10</td>
<td>95</td>
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<td>7</td>
<td>CH3CN</td>
<td>25</td>
<td>150</td>
<td>80</td>
</tr>
<tr>
<td>8</td>
<td>H2O</td>
<td>25</td>
<td>15</td>
<td>90</td>
</tr>
</tbody>
</table>

aIsolated yield. bReaction conditions: thiosemicarbazide 1 (X = S) (1 mmol), 1,3-diphenyl-1H-pyrazole-4-carbaldehyde 2a (1 mmol), solvent (3 mL), catalyst (0.06 gr, 4.7 mol%).

Table 2. Effect of catalyst types on the reaction time and yield of 3a

<table>
<thead>
<tr>
<th>Entry</th>
<th>Catalyst</th>
<th>Time (h)</th>
<th>Yield (%)a,b</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>-</td>
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<td>55</td>
</tr>
<tr>
<td>2</td>
<td>KSF</td>
<td>10</td>
<td>63</td>
</tr>
<tr>
<td>3</td>
<td>P-TSA</td>
<td>8</td>
<td>79</td>
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<tr>
<td>4</td>
<td>DABCO</td>
<td>4</td>
<td>82</td>
</tr>
<tr>
<td>5</td>
<td>nano-Fe3O4</td>
<td>3</td>
<td>87</td>
</tr>
<tr>
<td>6</td>
<td>CrCl3.6H2O</td>
<td>12</td>
<td>70</td>
</tr>
<tr>
<td>7</td>
<td>γ-Fe2O3@FAp</td>
<td>1</td>
<td>89</td>
</tr>
<tr>
<td>8</td>
<td>γ-Fe2O3@FAp@Cr</td>
<td>10 (min)</td>
<td>95</td>
</tr>
</tbody>
</table>

aIsolated yield. bReaction conditions: thiosemicarbazide 1 (X = S) (1 mmol), 1,3-diphenyl-1H-pyrazole-4-carbaldehyde 2a (1 mmol), CHCl3, room temperature, catalyst 4.7 mol%.

Table 3. Investigation of the amount of catalyst used in the synthesis of 3a

<table>
<thead>
<tr>
<th>Entry</th>
<th>amount of catalyst (g)</th>
<th>Time (min)</th>
<th>Yield (%)</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>0.05</td>
<td>20</td>
<td>85</td>
</tr>
<tr>
<td>2</td>
<td>0.06</td>
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<td>95</td>
</tr>
<tr>
<td>3</td>
<td>0.1</td>
<td>10</td>
<td>95</td>
</tr>
</tbody>
</table>

aIsolated yield. bReaction conditions: thiosemicarbazide 1 (X = S) (1 mmol), 1,3-diphenyl-1H-pyrazole-4-carbaldehyde 2a (1 mmol), CHCl3, room temperature.

This protocol was applied to the synthesis of a variety of pyrazole-triazole hybrids by using substituted pyrazole carbaldehydes (2) under the optimized reaction conditions and the results are presented in Table 4. This study reveals that both thiosemicarbazide (Entries 1-6) and semicarbazide (Entry 7) provide the desired products (Scheme 1) in high to excellent yields (88-95 %) and lower reaction times (10-15 min).
Table 4. Synthesis of pyrazole-triazole hybrids (3a-g) using γ-Fe₂O₃@FAP@Cr under optimized conditions.

<table>
<thead>
<tr>
<th>Entry</th>
<th>Product</th>
<th>Structure</th>
<th>Time (min)</th>
<th>M. p. Observed</th>
<th>M. p. Reported</th>
<th>Yield(^a) (%)</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>3a</td>
<td><img src="image" alt="Structure for 3a" /></td>
<td>10</td>
<td>226-228</td>
<td>225-227(^{40})</td>
<td>95</td>
</tr>
<tr>
<td>2</td>
<td>3b</td>
<td><img src="image" alt="Structure for 3b" /></td>
<td>10</td>
<td>208-210</td>
<td>208-210(^{40})</td>
<td>95</td>
</tr>
<tr>
<td>3</td>
<td>3c</td>
<td><img src="image" alt="Structure for 3c" /></td>
<td>12</td>
<td>180-183</td>
<td>182-184(^{40})</td>
<td>90</td>
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<tr>
<td>4</td>
<td>3d</td>
<td><img src="image" alt="Structure for 3d" /></td>
<td>15</td>
<td>219-221</td>
<td>218-220(^{40})</td>
<td>89</td>
</tr>
<tr>
<td>5</td>
<td>3e</td>
<td><img src="image" alt="Structure for 3e" /></td>
<td>10</td>
<td>248-250</td>
<td>247-250(^{40})</td>
<td>95</td>
</tr>
<tr>
<td>6</td>
<td>3f</td>
<td><img src="image" alt="Structure for 3f" /></td>
<td>10</td>
<td>235-238</td>
<td>234-236(^{40})</td>
<td>92</td>
</tr>
<tr>
<td>7</td>
<td>3g</td>
<td><img src="image" alt="Structure for 3g" /></td>
<td>15</td>
<td>180-182</td>
<td>180-182(^{40})</td>
<td>88</td>
</tr>
</tbody>
</table>

\(^a\)Isolated yield.
The synthesis of model compound (3a) in ethanol or methanol at room temperature provided 5-alkoxy-3-(3-(3-aryl)-1-phenyl-1\(H\)-pyrazol-4-yl)-4\(H\)-1,2,4-triazole derivatives (4a-e) (Scheme 2) in excellent yields (Table 5).

Scheme 2. Synthesis of 5-alkoxy-3-(3-(3-aryl)-1-phenyl-1\(H\)-pyrazol-4-yl)-4\(H\)-1,2,4-triazole derivatives via \(\gamma\)-Fe\(_2\)O\(_3\)@FAp@Cr.

Table 5. Synthesis of 5-alkoxy-3-(3-(3-aryl)-1-phenyl-1\(H\)-pyrazol-4-yl)-4\(H\)-1,2,4-triazole derivatives (4a-e) using \(\gamma\)-Fe\(_2\)O\(_3\)@FAp@Cr.

<table>
<thead>
<tr>
<th>Entry</th>
<th>Product</th>
<th>Structure</th>
<th>Time (min)</th>
<th>M. p. Measured</th>
<th>M. p. Reported</th>
<th>Yield(^a)</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>4a</td>
<td><img src="image1.png" alt="Structure" /></td>
<td>13</td>
<td>226-228</td>
<td>226-228(^{41})</td>
<td>90</td>
</tr>
<tr>
<td>2</td>
<td>4b</td>
<td><img src="image2.png" alt="Structure" /></td>
<td>10</td>
<td>215-217</td>
<td>216-218(^{41})</td>
<td>90</td>
</tr>
<tr>
<td>3</td>
<td>4c</td>
<td><img src="image3.png" alt="Structure" /></td>
<td>11</td>
<td>212-214(^{42})</td>
<td>214-216(^{41})</td>
<td>90</td>
</tr>
<tr>
<td>4</td>
<td>4d</td>
<td><img src="image4.png" alt="Structure" /></td>
<td>10</td>
<td>281-284</td>
<td>280-282(^{41})</td>
<td>95</td>
</tr>
<tr>
<td>5</td>
<td>4e</td>
<td><img src="image5.png" alt="Structure" /></td>
<td>10</td>
<td>&gt; 300</td>
<td>&gt; 300(^{41})</td>
<td>94</td>
</tr>
</tbody>
</table>

\(^a\)Isolated yield.
Interesting results obtained in the synthesis of pyrazole-triazole hybrids (Tables 4 & 5) encouraged us to extend the scope of this protocol to the synthesis of bis-triazole derivatives (7a-c) (Scheme 3). Initially, the reaction of 1,4-dibromobutane with thiosemicarbazide resulted in the synthesis of butane-1,4-diyl-bis(hydrazinecarbimidothioate) which then reacted with various aromatic aldehydes to furnish the novel derivatives of bis-triazoles in excellent yields and reasonable reaction times (Table 6). Structure of these new bis-triazole derivatives were established by spectroscopic (FT-IR, ¹H NMR, ¹³C NMR, Mass) analyses.

Scheme 3. Synthesis of novel derivatives of bis-triazoles using γ-Fe₂O₃@FAp@Cr.

Table 6. Synthesis of novel derivatives of bis-triazoles using γ-Fe₂O₃@FAp@Cr at room temperature.

<table>
<thead>
<tr>
<th>Entry</th>
<th>product</th>
<th>Structure</th>
<th>Time (min)</th>
<th>M. p. (°C)</th>
<th>Yielda (%)</th>
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<tbody>
<tr>
<td>1</td>
<td>7a</td>
<td><img src="image" alt="Structure of 7a" /></td>
<td>60</td>
<td>&gt; 300</td>
<td>95</td>
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<tr>
<td>2</td>
<td>7b</td>
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<td>70</td>
<td>&gt; 300</td>
<td>90</td>
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<td>65</td>
<td>&gt; 300</td>
<td>92</td>
</tr>
</tbody>
</table>

aIsolated yield.
A suggested mechanism for the synthesis of triazole derivatives in the presence of chromium incorporated fluorapatite encapsulated iron oxide nanocatalyst (γ-Fe₂O₃@FAP@Cr) is presented in Scheme 4. Initially, γ-Fe₂O₃@FAP@Cr NPs activate pyrazolecarbaldehyde via coordination to the carbonyl group of the aldehyde. In continuation, thiosemicarbazide or semicarbazide is added to the activated carbonyl group of aldehyde producing arylidene intermediates A or B which by intermolecular cyclization furnish the target products 3a-g, 4a-e and 7a-c.

Scheme 5. A proposed mechanism for the synthesis of triazole derivatives using γ-Fe₂O₃@FAP@Cr nanocatalyst.

The recyclability of the catalyst was investigated in the synthesis of model compound 3a. At the end of each reaction the nanocatalyst was separated by an external magnet, washed with hot ethanol, dried at 80°C and reused in the subsequent run. This study showed that after six consecutive cycles the catalytic activity was preserved without any striking loss in its catalytic activities (Fig.7).

3. Conclusions

We have introduced γ-Fe₂O₃@FAP@Cr, as an efficient, novel, inexpensive, eco-friendly and recyclable nanocatalyst, for the synthesis of pyrazole-triazole hybrids and bis-triazoles. The prominent advantages of this method can be described as adherence to the basis of green chemistry, easy work-up procedure without any need for chromatographic separation, short reaction times, excellent yields, facile removal and reuse of the catalyst.
Fig. 7. Recyclability of $\gamma$-Fe$_2$O$_3$@FAP@Cr in the synthesis of 3a as the model compound.

Acknowledgements

Partial financial support of University of Guilan for this research work is sincerely acknowledged.

4. Experimental

4.1. Materials and Methods

Chemicals for this research were purchased from Merck and Fluka. Melting points were determined on a Böchi B-545 apparatus in open capillary tubes. FT-IR spectra were recorded on a $\alpha$-Bruker spectrometer. $^1$H NMR spectra were recorded on a 300 MHz Bruker DRX-300 in DMSO-d$_6$ as solvent and tetramethylsilane (TMS) as internal standard. $^{13}$C NMR spectra were obtained on a 75 MHz Bruker DRX-75 in DMSO-d$_6$ as solvent. Mass spectra were obtained from AB SCIEX 3200 QTRAP. XRD was done on a KEFA Analytical XPERT-PRO. Scanning Electron Microscope (SEM) were investigated on a model: VP 1450, company: LEO-Germany. Elemental analysis (EDX) was obtained on Oxford Instruments EDS Microanalysis X-MAX-80; model: TeScan-Mira III. Transmission electron microscopy (TEM) measurements were recorded on a Zeiss-EM10C-100 KV instrument. Thin layer chromatography (TLC) was done with ethyl acetate: n-hexane 1:1 on TLC Silica gel 60 F$_{254}$.

4.2. Synthesis of $\gamma$-Fe$_2$O$_3$@FAP@Cr MNPs

$\gamma$-Fe$_2$O$_3$@FAP MNPs was prepared according to the reports$^{37,38}$. 125 mg of $\gamma$-Fe$_2$O$_3$@FAP was stirred with 2 mmol CrCl$_3$.6H$_2$O in 25 ml water at room temperature for a period of 1 h. The obtained slurry was magnetic decanted, washed with DW frequently, and dried at 100 °C to give $\gamma$-Fe$_2$O$_3$@FAP@Cr NPs as a brown solid (655 mg).

4.3. General procedure for the synthesis of 5-(3-(aryl)-1-phenyl-1H-pyrazol-4-yl)-1,2,4-triazolidine-3-ones (thiones) and 5-alkoxy-3-(3-(3-aryl)-1-phenyl-1H-pyrazol-4-yl)-4H-1,2,4-triazole derivatives

A mixture of semicarbazide or thiosemicarbazide (1 mmol), aromatic aldehyde (1 mmol) and 0.06 g $\gamma$-Fe$_2$O$_3$@FAP@Cr were stirred in chloroform (5 mL) at room temperature for the required reaction time and the progress of the reaction was monitored by thin layer chromatography (ethyl acetate: n-hexane 1: 2). After completion of the reaction, $\gamma$-Fe$_2$O$_3$@FAP@Cr was separated by an external magnet
(1.4 Tesla) and washed with hot DW and ethanol three times, dried and reused in the next run under similar reaction conditions. The reaction mixture after separation of the catalyst evaporated under vacuum and the residue was recrystallized from ethanol to produce the desired pyrazole-triazole products (Table 4). The reaction in methanol or ethanol as a solvent gave 5-alkoxy-3-(3-(3-aryl)-1-phenyl-1H-pyrazol-4-yl)-4H-1,2,4-triazoles presented in Table 5.

4.4. preparation of butane-1,4-diyl-bis(hydrazinecarbimidothioate)

A mixture of 1,4-dibromobutane (1 mmol) and thiosemicarbazide (2 mmol) were stirred at room temperature in ethanol (5 mL) and the progress of the reaction was checked by TLC (ethyl acetate: n-hexane 8: 2). After completion of the reaction, butane-1,4-diyl-bis(hydrazinecarbimidothioate) was obtained as a milky solid with 96% yield.

4.6. General procedure for the synthesis of 1,4-bis((3-Aryl-1H-1,2,4-triazol-5-yl)thio)butane

A mixture of arylaldehyde (2 mmol) and butane-1,4-diyl-bis(hydrazinecarbimidothioate) (1 mmol) and 0.06 g γ-Fe2O3@FAp@Cr were stirred at room temperature for the required reaction time (Table 6) and the progress of the reaction was monitored by thin layer chromatography (ethyl acetate: n-hexane 1: 1). After completion of the reaction, γ-Fe2O3@FAp@Cr nanoparticles were separated by an external magnet (1.4 Tesla) and washed with hot ethanol three times, dried and reused in the next run under similar reaction conditions. After removal of the catalyst the reaction mixture was evaporated under vacuum by a rotary evaporator. The resulting solid residue was purified by recrystallization from ethanol to produce 1,4-bis((3-aryl-1H-1,2,4-triazol-5-yl)thio)butane in 90-95 % yield.

4.7 Physical and spectral data of selected compounds

1,4-bis((3-(3-(4-nitrophenyl)-1-phenyl-1H-pyrazol-4-yl)-1H-1,2,4-triazol-5-yl)thio)butane (7a). Yield 95%, light brown solid; M. p. 300 > °C. FT-IR (KBr), \( \nu \), cm\(^{-1}\): 3309 (N-H), 3094, 2976, 2892 (C-H), 1663, 1612, 1517 (C=C and C=N), 1563, 1379 (NO\(_2\)), 1253 (C-S-C), 827, 786, 696. 1H NMR (300 MHz, DMSO-d\(_6\)), \( \delta \), ppm: 1.83 (br. s, 2H, CH\(_2-\)CH\(_2-S\)), 3.38 (br. s, 2H, CH\(_2-S\)), 7.50 (t, \( J = 7.2 \) Hz, 1H), 7.61 (t, \( J = 7.5 \) Hz, 2H), 7.94 (d, \( J = 7.5 \) Hz, 2H), 7.99 (d, \( J = 8.7 \) Hz, 2H), 8.33 (s, 1H, CH-N), 8.37 (d, \( J = 8.7 \) Hz, 2H), 9.23 (s, 1H, NH). 13C NMR spectrum (75 MHz, DMSO-d\(_6\)), \( \delta \), ppm: 156.4 (S-C-NH), 151.4 (C-C-NH), 141.1, 138.6, 130.2, 129.3, 129.2, 128.7, 128.5, 128.3, 128.1, 120.1, 115.1, 114.0, 113.9, 29.7 (CH\(_2-S\)), 26.2 (CH\(_2-\)CH\(_2-S\)). MS, \( m/z \): 736.7 (M-NO\(_2\)), 518.8 (M-C\(_{15}H\_10N_3O_2\)). Anal. Calcd. For C\(_{38}H\_30N_12O_4S_2\) (782.2): C, 58.30; H, 3.86; N, 21.47. Found: C, 58.22; H, 3.67; N, 21.29.

1,4-bis((3-(1,3-diphenyl-1H-pyrazol-4-yl)-1H-1,2,4-triazol-5-yl)thio)butane (7b). Yield 90%, cream solid; M. p. 300 > °C. FT-IR (KBr), \( \nu \), cm\(^{-1}\): 3248 (N-H), 3026 (C-H), 1656, 1587, 1535, 1460 (C=C and C=N), 1384, 1271 (C=C-NH), 1236, 118.9, 109.8, 29.9 (CH\(_2-S\)), 26.2 (CH\(_2-\)CH\(_2-S\)). MS, \( m/z \): 736.7 (M-NO\(_2\)). Anal. Calcd. For C\(_{38}H\_32N_10S_2\) (692.9): C, 65.87; H, 4.66; N, 20.22. Found, (%): C, 65.75; H, 4.55; N, 20.08.

1,4-bis((3-(4-chlorophenyl)-1-phenyl-1H-pyrazol-4-yl)-1H-1,2,4-triazol-5-yl)thio)butane (7c). Yield 92%, cream solid; M. p. 300 > °C. FT-IR (KBr), \( \nu \), cm\(^{-1}\): 3426, 3249 (N-H), 2923, 2873 (C-H), 1535, 1460 (C=C and C=N), 1384, 1271, 1020 (C-C-S-C), 856, 779, 728. 1H NMR spectrum (300 MHz, DMSO-d\(_6\)), \( \delta \), ppm: 1.77 (br. s, 2H, CH\(_2-S\)), 3.36 (br. s, 2H, CH\(_2-S\)), 7.54 (t, \( J = 7.4 \) Hz, 1H), 7.54-7.62 (m, 6H), 7.82 (s, 1H, CH-N), 7.82 (d, \( J = 8.1 \) Hz, 2H), 9.08 (s, 1H, NH). 13C NMR spectrum (75 MHz, DMSO-d\(_6\)), \( \delta \), ppm: 163.7 (S-C-NH), 138.3, 134.3, 130.2, 129.3, 129.1, 129.1, 128.9, 128.7, 126.5, 123.6, 118.2, 109.8, 29.9 (CH\(_2-S\)), 27.2 (CH\(_2\)-CH\(_2-S\)). MS, \( m/z \): 507.5 (M-2C\(_{15}H\_10ClN_2\)). Anal. Calcd. For C\(_{38}H\_30Cl_2N_10S_2\) (761.75): C, 59.92; H, 3.97; N, 18.21. Found: C, 59.92; H, 3.97; N, 18.39.
References

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